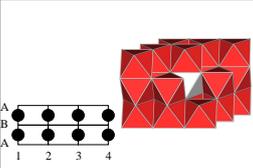
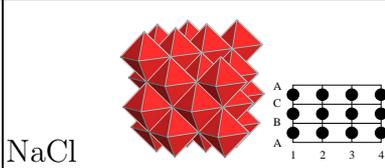
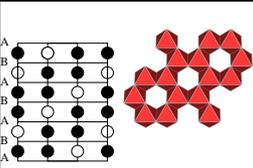
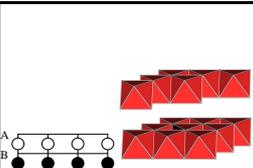
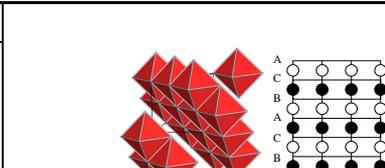
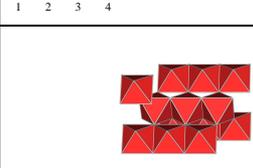
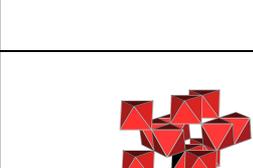
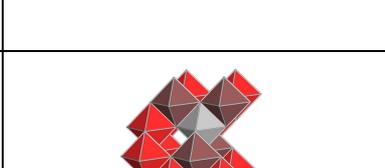
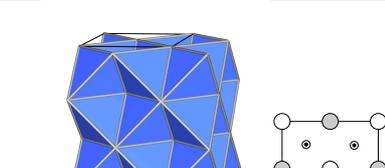
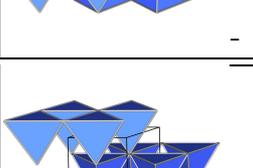
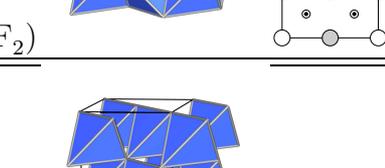


## Strukturchemie von Metallen und Ionenkristallen II

## Dichte Packungen mit gefüllten Oktaederlücken

$X$ hexagonal dicht	Formel	CN ( $X$ )	Lücken- besetzung	$X$ kubisch dicht
 NiAs	$MX$	6	1:	 NaCl
 Korund	$M_2X_3$	4	2/3:	-
 CdI <sub>2</sub>	$MX_2$	3	1/2	 CdCl <sub>2</sub>
 CaCl <sub>2</sub> (Rutil)	Raum-		1/2:	
 $\alpha$ -PbO <sub>2</sub>	netz		1/2:	 Anatas (TiO <sub>2</sub> )

## Dichte Packungen mit gefüllten Tetraederlücken

$X$ hexagonal dicht	Formel	CN ( $X$ )	Lücken- bes.	$X$ kubisch dicht
 -	$M_2X$	4:8	1	 Li <sub>2</sub> O (a-CaF <sub>2</sub> )
 Wurtzit	$MX$	4:4	1/2	 Zinkblende